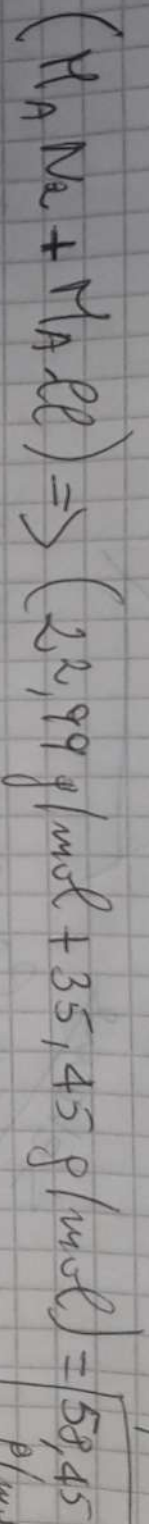


2 gr

$$\text{mol} = \frac{m(g)}{M(g/mol)} = \frac{2g}{58,45} = 0,034$$



nombre molar atomique



1

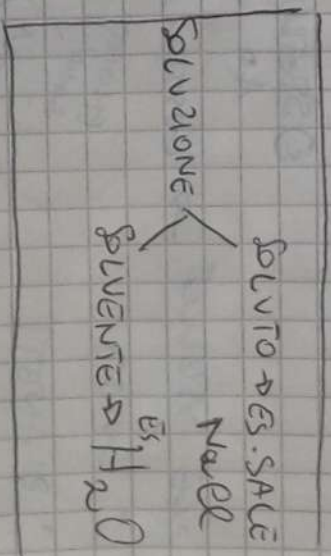
1 mol: $6,022 \cdot 10^{23} = 0,034 \text{ mol} : \textcircled{X}$ portable

$X = \frac{6,022 \cdot 10^{23} \cdot 0,034 \text{ mol}}{1 \text{ mol}} =$

1

$C = \frac{\text{QUANTITE SOLUTO}}{\text{QUANTITE SOLVENTO}}$

(L) QUANTITE SOLVENTO



$\% m/m = \frac{m \text{ soluto}}{m \text{ soluto} + m \text{ solvante}} \cdot 100$